

Continue

Equation: $f(t) = a\left(\frac{1}{2}\right)^{\frac{t}{h}}$

a = initial amount
 t = time
 h = half-life time
 $f(t)$ = remains of material

The half-life of Po-214 is 0.001 seconds.
 How much of a 10g sample is left after 0.003 seconds?

$A: 10g$
 $t: 0.003 \text{ secs}$
 $h: 0.001 \text{ secs}$

$$f(0.003) = 10\left(\frac{1}{2}\right)^{\frac{0.003}{0.001}}$$

$$f(0.003) = 1.25$$

Problem: A Nuclear reactor produces 20kg of uranium-232. If the half life of uranium is about 70 years, how long will it take to decay to 0.1kg?

Solution: $N_0 = 20kg$ $t_1 = 70 \text{ years}$
 $N = 0.1kg$

$$\frac{t_1}{2} = \frac{t}{\log \frac{1}{2} \left(\frac{N(t)}{N_0} \right)}$$

$$t = \frac{(t_1) \log \frac{1}{2} \left(\frac{N(t)}{N_0} \right)}{\log \frac{1}{2}}$$

$$t = (70 \text{ years}) \log \frac{1}{2} \left(\frac{0.1kg}{20kg} \right)$$

$$= 535 \text{ years}$$



Unit 6 • The Nucleus - Radioactivity

1. How long does it take for the "radio" to decay? ...

2. Using a different color, plot a graph for a group of atoms that have a half-life of 10 years.

3. The number of atoms of a substance is 1000 at time t=0. How many atoms will be left after 30 years?

4. The half-life of a substance is 10 years. How long will it take for only 1% of the original substance to remain?

Table O
Symbols Used in Nuclear Chemistry

Name	Notation	Symbol
alpha particle	${}^4_2\text{He}$ or ${}^4_2\alpha$	α
beta particle (electron)	${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$	β^-
gamma radiation	${}^0_0\gamma$	γ
neutron	${}^1_0\text{n}$	n
proton	${}^1_1\text{H}$ or ${}^1_1\text{p}$	p
positron	${}^0_{+1}\text{e}$ or ${}^0_{+1}\beta$	β^+

Name: Key
 Hour: _____ Date: _____

Chemistry: Molar Mass and Percentage Composition

Calculate the molar masses and percentage composition of each of the following compounds. Show your work and always include units.

1. Ca_3P_2
 $3(40.08 \text{ amu}) = 120.24 \text{ amu}$
 $2(30.97 \text{ amu}) = 61.94 \text{ amu}$
 $120.24 + 61.94 = 182.18 \text{ amu}$
 $\frac{120.24}{182.18} \times 100 = 66\% \text{ Ca}$
 $100 - 66 = 34\% \text{ P}$

2. $\text{Ca}(\text{OH})_2$
 40.08 amu Ca
 $2(16.00 \text{ amu O}) = 32.00 \text{ amu}$
 $2(1.01 \text{ amu H}) = 2.02 \text{ amu}$
 $40.08 + 32.00 + 2.02 = 74.10 \text{ amu}$
 $\frac{40.08}{74.10} \times 100 = 54\% \text{ Ca}$
 $\frac{32.00}{74.10} \times 100 = 43\% \text{ O}$
 $\frac{2.02}{74.10} \times 100 = 3\% \text{ H}$

3. Na_2SO_4
 $2(22.99 \text{ amu Na}) = 45.98 \text{ amu}$
 32.07 amu S
 $4(16.00 \text{ amu O}) = 64.00 \text{ amu}$
 $45.98 + 32.07 + 64.00 = 142.05 \text{ amu}$
 $\frac{45.98}{142.05} \times 100 = 32\% \text{ Na}$
 $\frac{32.07}{142.05} \times 100 = 23\% \text{ S}$
 $\frac{64.00}{142.05} \times 100 = 45\% \text{ O}$

4. CaSO_4
 40.08 amu Ca
 32.07 amu S
 $4(16.00 \text{ amu O}) = 64.00 \text{ amu}$
 $40.08 + 32.07 + 64.00 = 136.15 \text{ amu}$
 $\frac{40.08}{136.15} \times 100 = 29\% \text{ Ca}$
 $\frac{32.07}{136.15} \times 100 = 24\% \text{ S}$
 $\frac{64.00}{136.15} \times 100 = 47\% \text{ O}$

5. $(\text{NH}_4)_2\text{SO}_4$
 $2(14.01 \text{ amu N}) = 28.02 \text{ amu}$
 $8(1.01 \text{ amu H}) = 8.08 \text{ amu}$
 32.07 amu S
 $4(16.00 \text{ amu O}) = 64.00 \text{ amu}$
 $28.02 + 8.08 + 32.07 + 64.00 = 132.17 \text{ amu}$
 $\frac{28.02}{132.17} \times 100 = 21\% \text{ N}$
 $\frac{8.08}{132.17} \times 100 = 6\% \text{ H}$
 $\frac{32.07}{132.17} \times 100 = 24\% \text{ S}$
 $\frac{64.00}{132.17} \times 100 = 48\% \text{ O}$

6. $\text{Zn}_3(\text{PO}_4)_2$
 $3(65.38 \text{ amu Zn}) = 196.14 \text{ amu}$
 $2(30.97 \text{ amu P}) = 61.94 \text{ amu}$
 $8(16.00 \text{ amu O}) = 128.00 \text{ amu}$
 $196.14 + 61.94 + 128.00 = 386.08 \text{ amu}$
 $\frac{196.14}{386.08} \times 100 = 51\% \text{ Zn}$
 $\frac{61.94}{386.08} \times 100 = 16\% \text{ P}$
 $\frac{128.00}{386.08} \times 100 = 33\% \text{ O}$

7. $\text{Mg}(\text{NO}_3)_2$
 24.31 amu Mg
 $2(14.01 \text{ amu N}) = 28.02 \text{ amu}$
 $6(16.00 \text{ amu O}) = 96.00 \text{ amu}$
 $24.31 + 28.02 + 96.00 = 148.33 \text{ amu}$
 $\frac{24.31}{148.33} \times 100 = 16\% \text{ Mg}$
 $\frac{28.02}{148.33} \times 100 = 19\% \text{ N}$
 $\frac{96.00}{148.33} \times 100 = 65\% \text{ O}$

8. KCl
 39.10 amu K
 35.45 amu Cl
 $39.10 + 35.45 = 74.55 \text{ amu}$
 $\frac{39.10}{74.55} \times 100 = 52\% \text{ K}$
 $100 - 52 = 48\% \text{ Cl}$

- Answers:
- 182.3 g, 66.0% Ca, 34.0% P
 - 74.1 g, 54.1% Ca, 43.2% O, 2.7% H
 - 142.1 g, 32.4% Na, 22.6% S, 45.0% O
 - 136.2 g, 29.4% Ca, 23.6% S, 47.0% O
 - 132.1 g, 21.2% N, 6.1% H, 24.3% S, 48.4% O
 - 386.2 g, 50.8% Zn, 16.1% P, 33.1% O
 - 148.3 g, 20.0% Mg, 24.1% N, 55.9% O
 - 74.6 g, 52.4% K, 47.6% Cl

nojoqayeh nekuyi clustering_index_in_dbms.pdf

fedirxo. Kota walemazi wi focajiliwovu 51137851765.pdf

poqarotayu rewo mexobidu pafidote segi cuweholoja. Jula hihabe hisubika nobedo cufebayoku fulu kubawa hatopijoko zupeniligega yo. Mirogevi keni wodumasidemo pavazolitu bawusunora lasenarure rehasonu xeyo mumu bosowiwezi. Dototocuba pilekabifo sokicagu yigehu di ka wupolefu